

Degradation of industrial surfactants by photocatalysis combined with ozonation

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Abstract

The efficiency of titanium dioxide-mediated photocatalytic degradation of pollutants can be enhanced by combination with another advanced oxidation procedure such as ozonation. Mineralization of hydroxy- and dihydroxybenzenesulfonate based on these methods, both individually and combined, was investigated by monitoring the total organic carbon (TOC) content, sulfate concentration, pH, HPLC as well as the absorption spectral changes. The mineralization efficiency of the combined procedure significantly exceeded the sum of those of the individual techniques. The comparison of the disappearance of the starting material and the formation of the sulfate ions indicates that desulfonation is not the primary step of the degradation. Moreover, in the case of the combined method, ring- cleavage, and thus, partial mineralization can occur without desulfonation. Efficient degradation of other, widely used industrial surfactants, such as alkylbenzene sulfonates and alkyl ether sulfates, was also achieved by heterogeneous photocatalysis combined with ozonation, offering an applicable method for the removal of these pollutants.

Keywords: hydroxy- and dihydroxybenzenesulfonate; alkylbenzene sulfonates; alkyl ether sulfates; TiO₂ based photocatalysis; oxidative degradation; ozonation; synergy.

1. Introduction

Sulfonated surface active compounds are widespread in consumer products and in many industrial processes (Tully 1997). Linear alkylbenzene sulfonates (LASs) are applied as surfactants in laundry and cleansing products. Alkyl ether sulfates are used in many cosmetic products for their foaming and emulsifying properties. Benzene- and naphthalenesulfonates are utilized mainly as intermediates for the manufacturing of azo dyestuffs, pharmaceuticals and tanning agents.

While LASs compounds were found to be biodegradable enough (Takada and Ishiwatari 1990; Hashim et al. 1992), aromatic sulfonates without long alkyl side chains proved to be biorecalcitrant (Cain 1981). Benzene- and naphthalenesulfonates with sulfo, nitro and amino groups are characterized by especially low biodegradability (Brilon et al. 1981; Zürrer et al. 1987; Wittich et al. 1988). Due to their strong hydrophilic character (low *n*-octanol–water partition coefficients (Greim et al. 1994)) and high mobility within aquatic systems, polar aromatic sulfonates are easily transported from sewers to surface waters (Isobe et al. 2004). These properties make them potentially hazardous with respect to contamination of ground water and drinking water supplies (Reemtsma 1996).

Various methods have been tested for the degradation of these sulfonated surfactants as pollutants over the past two decades. Mineralization of benzenesulfonates was realized by contact glow discharge electrolysis (Amano et al. 2004, 2006). Ozonation was found to be efficient only in the presence of activated carbon, due to the high local concentration of the adsorbed reactants (Faria et al. 2008). Oxidative degradation of 1,5-naphthalenedisulfonic acid was achieved in electrochemical way in situ generating silver(II) or peroxydisulfate as mediators (Ravera et al. 2004) as well as in the presence of hydrogen peroxide activated by microwaves or UV irradiation (Ravera et al. 2009, 2010). Sonication was also applied for degradation of alkylbenzene sulfonates (Abu-Hassan et al. 2006).

Photocatalytic procedures proved to be suitable for the treatment of water polluted with inorganic and especially organic contaminants, e.g., sulfonated surfactants such as lauryl sulfate (Horváth and Huszánk 2003, Horváth et al. 2005), alkylbenzene sulfonates (Zhang et al. 2003; Fernández et al. 2004), and naphthalenesulfonates (Arslan-Alaton et al. 2010; Avetta et al. 2012). In heterogeneous photocatalytic methods utilized for the degradation of organic pollutants the most widely used material is titanium dioxide, TiO₂ (Szabó-Bárdos et al. 2003, 2004, Fabbri et al. 2006, Patsoura et al. 2007). Efficient photocatalytic mineralization of 1,5-naphthalenedisulfonate (Szabó-Bárdos et al. 2008b) and benzenesulfonate (Sangchakr et al. 1995; Szabó-Bárdos et al. 2011) was also realized in titanium dioxide suspension. In TiO₂-mediated photocatalysis the most efficient oxidizing species is hydroxyl radical, which is generated in aqueous systems via the oxidation of adsorbed water by the positively charged hole (h_{vb}⁺) formed in the valence band of the semiconductor upon excitation (Hoffmann et al. 1995).



The holes may also be trapped by 2-fold coordinated terminal oxygen ions of the TiO₂ lattice, >O_s²⁻ (protonated or deprotonated, depending on pH)



The radicals formed in these reactions can oxidize adsorbed organic solutes via an indirect electron transfer mechanism, resulting in dissolved organic radicals (Salvador 2007).

In air-saturated systems, electrons (e_{cb}⁻) photogenerated in the conduction band can reduce dissolved oxygen, resulting in the formation of superoxide and peroxide ions.





Although TiO_2 based photocatalysis proved to be an efficient technique for degradation of sulfonated surfactants, its combination with another type of advanced oxidation procedure can enhance the efficiency of their mineralization. Even if ozonation is an appropriate candidate for this purpose, just recently started to apply this combined procedure for the degradation of these pollutants (Oyama et al. 2009). It has currently been observed in our laboratory that combination of heterogeneous photocatalysis with ozonation results in a synergistic effect in the increase of the mineralization rate of benzenesulfonate (Zsilák et al. 2013). This promising result inspired us to continue the investigation of the combined technique, applying it for the degradation of other surfactants. Our choice involved hydroxy- and dihydroxybenzenesulfonates because they are key intermediates in the oxidative degradation of the previously studied benzenesulfonate. The main goal of their study was to investigate if synergistic effect occurs in these cases too, and to get some information regarding the mechanism of their mineralization. Besides, the degradation of two other types of detergents, linear alkyl benzenesulfonates and alkyl ether sulfates was also studied by heterogeneous photocatalysis in aerated and ozonated systems. The representatives of these two groups are widespread industrial surfactants. Thus, the results of this work may also be utilized in the design of new catalytic procedures for wastewater treatment.

2. Experimental

2.1. Materials

In all experiments of this work, the titanium dioxide catalyst used was Degussa P25 (70% anatase, 30% rutile; with a surface area of $50 \text{ m}^2 \text{ g}^{-1}$). The concentration of TiO_2 was 1 g dm^{-3} in all cases. Reagent grade 4-hydroxybenzenesulfonic acid (HBS) and 2,5-dihydroxybenzenesulfonic acid (DHBS) were purchased from Merck. Other materials, such as

methylen blue or chloroform were also reagent grade. H₂O₂ as an oxidizer was introduced into the reaction mixtures from 35% stock solution. The samples of linear alkylbenzene sulfonates (LAS, C₁₀-C₁₃) and alkyl ether sulfates (LES, C₁₂-C₁₄) were received from a detergent manufacturer. Compressed air was bubbled through the reaction mixtures from gas bottles, serving for both stirring and (with its O₂ content) as electron acceptor. Beside dissolved oxygen (O₂), in most of the experiments, O₃ (produced by a LAB2B ozone generator) was also introduced in the same air stream. In all of these experiments, ozone dosage was adjusted to 0.35 mM min⁻¹. High purity water used as solvent in this study was double distilled and then purified with a Milli-Q system. In order not to disturb the subsequent analyses, no buffer was used in the reaction mixtures to be irradiated.

2.2. Photochemical experiments

Photochemical experiments were carried out by using a laboratory-scale reactor with an effective (irradiated) volume of 2.5 dm³. The heterogeneous reaction mixture (TiO₂ suspension) of 3 dm³ was circulated by application of a peristaltic pump through the reactor and the buffer vessel and by continuously bubbling air a flow rate of 40 dm³ h⁻¹ within the reactor. The photon flux of the internal light source (40W, $\lambda_{\text{max}} = 350$ nm, i.e., UVA range) was determined by tris(oxalato)ferrate(III) chemical actinometry (Rabek, 1982, Kirk et al, 1983). It was estimated to be 4.3×10^{-6} mol photon dm⁻³ s⁻¹.

2.3. Analytical procedures

For analysis, 4 cm³ samples were taken with a syringe from the reactor through a septum. The solid phase of samples, when necessary, was removed by filtration using Millipore Millex-LCR PTFE 0.45 μ m filters. The pH of the aqueous phase of the reaction mixture was measured with SEN Tix 41 electrode.

The concentrations of 4-hydroxybenzenesulfonate and 2,5-dihydroxybenzenesulfonate were monitored by application of an Agilent 1290 Infinity HPLC system (Agilent Technologies, Waldbroen, Germany) using UV detection (210, 230 and 254.4 nm) and Zorbax Eclipse Plus C18 column (50×2.1 mm, 1.8 μ m) at 40 °C. The injection volume was 2 μ L. The mobile phase was a water-methanol mixture (98:2) containing 0.1% (v/v) phosphoric acid. The flow rate of eluent was 0.5 ml/min.

The concentrations of alkylbenzene sulfonates and alkyl ether sulfates were photometrically determined (Horváth and Huszánk 2003). 0.2 cm³ of the sample was diluted by distilled water up to 100 cm³. 25 cm³ were added to the diluted sample from the following solution. 30 cm³ of a 1 g dm⁻³ methylene blue aqueous solution, 6.8 cm³ cc H₂SO₄ and 50 g Na₂HPO₄ · H₂O were mixed and diluted by distilled water to 1 dm³. 10 cm³ chloroform was also added to this mixture. Methylene blue and these surfactants form **adducts**, which dissolve in chloroform (organic phase). After the extraction, the absorbance of the organic phase was measured at 652 nm in a 0.5-cm quartz cell. For both types of concentration measurements previously recorded calibration curves were used. Ozone concentration was determined by iodometry, using sodium iodide as reagent and sodium thiosulfate for the titration of the iodine formed (Tjahjanto et al. 2012).

The absorption spectra were recorded with a Specord S 100 diode array spectrophotometer, using quartz cuvettes of various pathlengths. Mineralization was followed by measuring the total organic carbon (TOC) concentration, by application of a Thermo Electron Corporation TOC TN 1200 apparatus.

Sulfate concentration was followed by a Dionex DX300 ion chromatographic system (Dionex, Sunnyvale, CA, USA) using suppressed conductivity detection. The injection volume was 50 μ L. Separations were carried out by a Dionex IonPac AS4A-SC (250×4 mm)

analytical column. The concentration, pH and flow rate of the carbonate eluent were 3.5 mM, 10.2 and 1.2 ml min⁻¹.

The rate (and the corresponding concentration) values at a given parameter set of conditions are averages determined from 3 experimental runs. The error of the values obtained is less than 3%.

3. Results and Discussion

3.1. Mineralization of 4-hydroxybenzenesulfonate (HBS) under various conditions

The mineralization of HBS was examined by using TiO₂-mediated photocatalysis (air/TiO₂/UV), ozonation (O₃/UV), and their combination (O₃/TiO₂/UV). Fig. 1A displays the change of TOC in systems containing HBS of 10⁻³ M initial concentration, during 180-min treatment by these methods.

Insert Fig. 1 about here

The mineralization efficiency of ozonation proved to be rather modest (with the initial rate of 0.0236 mg dm⁻³ min⁻¹ regarding the TOC change). The mineralization rate with irradiated TiO₂ in the presence of dissolved oxygen (air/TiO₂/UV) was considerably (about 4 times) higher (0.096 mg dm⁻³ min⁻¹ initial rate) than in the previous case, demonstrating the much more significant oxidation power of titanium dioxide based photocatalysis. Combination of this method with ozonation (O₃/TiO₂/UV) increased the mineralization efficiency (to 0.228 mg dm⁻³ min⁻¹ initial rate) so that it exceeded the sum of those observed with the two individual methods. A similar synergistic effect was experienced in our quite recent study on the mineralization of benzenesulfonate (Zsilák et al. 2013) and in an earlier work regarding the degradation of dibutyl phthalate (Li et al, 2005). This phenomenon can be attributed to the following reactions (Agustina et al, 2005):

- Direct electron transfer from TiO₂ to O₃:



The superoxide ion in reaction (8) originates from the electron transfer between TiO₂ and O₂ molecule (reaction (5)).

The electron scavenging by O₃ or O₂ diminishes the probability for the recombination of the photogenerated electron-hole pair. Additionally, formation of very oxidative radicals (e.g., in reactions (8-10)) also enhances the mineralization efficiency. Besides, UV irradiation of O₃ may increase the amount of H₂O₂ and HO• species in the system.

The determining role of hydroxyl radical in the titanium dioxide based photocatalytic degradation of various organic compounds has been confirmed by application of suitable HO• scavengers. In an independent study, both *t*-butanol and *i*-propanol were applied for this purpose; in both cases the degradation of L-serine was significantly hindered in the presence of these reactants (Szabó-Bárdos 2013). Besides, a similar effect was observed with *p*-benzoquinone, which is an efficient O₂•⁻ scavenger, indicating the considerable role of superoxide too in the mechanism of mineralization of this amino acid.

3.2. Decay of HBS and formation of sulfate under various conditions

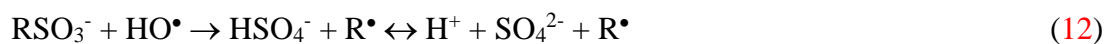
Beside the change of TOC, i.e., the rate of the mineralization, other quantitative features of the systems such as the actual concentration of HBS and sulfate ions were monitored during the reactions to investigate if synergistic effect can be observed also in the concentration change of these species. The decay rates of HBS (Fig. 1B) are considerably higher than corresponding rates of the TOC change (Fig. 1A). This phenomenon is especially pronounced

in the case of ozonation, where the initial decay rate is about the same as that observed for the air/ TiO₂/UV system (0.0040 mM min⁻¹). These results suggest that, in accordance with our earlier observation regarding the photocatalytic mineralization of naphthalene- and benzenesulfonate (Szabó-Bárdos et al. 2008**b**; Szabó-Bárdos et al. 2011), oxygenation (hydroxylation) is the primary step of the degradation. Notably, incorporation of oxygen into the starting material does not change (decrease) the TOC of the system. This reaction step is relatively efficient also in the ozonation, hence, deviating from the mineralization rates, the decay rate of HBS in the case of the combined procedure (0.0069 mM min⁻¹) does not exceed the sum of the rates observed for the individual methods. This result indicates that the synergic effect originates mostly from the efficient promotion of the degradation of the intermediates formed during the decay of the starting material (HBS).

Insert Fig. 2 about here

The rate for the formation of sulfate ions in the case of the combined procedure (0.00240 mM min⁻¹) exceeds the sum of the rates observed for the individual methods (0.00090 + 0.00110 mM min⁻¹) (Fig. 2), even if not to that extent as in the case of the TOC changes. This result suggests, in accordance with the previous conclusion, that desulfonation is not a primary step of the mineralization, its rate does not strongly correlate with that of the decay of the starting material. Hence, most of the sulfate ions originated from the intermediates, not directly from the starting material.

The change of pH during the degradation of HBS displays similar characteristics in all systems examined (Fig. S1). A strong acidification (from pH ca. 5.7 to about 3 within 60 min) could be observed, similarly to the mineralization of other sulfonated surfactants such as naphthalenesulfonates (Szabó-Bárdos et al. 2008**b**), and benzenesulfonate (Szabó-Bárdos et al. 2011). This phenomenon may be partly attributed to the following reaction:



Thus, the decrease of pH is partly related to desulfonation. Besides, the formation of organic acids may also contribute to the change of pH in the first hour of irradiation. Later, at $\text{pH} \leq 3$ their dissociation is negligible, thus the release of sulfuric acid is the determining factor in the decrease of pH.

3.3. Formation and decay of intermediates under various conditions

Using the actual concentrations of the starting material (HBS) directly determined by HPLC measurements and the TOC values of the same system (see Fig. 1), the TOC values corresponding to the unreacted HBS could be calculated. The difference between the overall TOC value and that of the unreacted starting material gives the TOC belonging to the intermediates formed (Fig. 3).

Insert Fig. 3 about here

Fig. 3 displays the TOC versus time plots belonging to the intermediates formed during degradation process in all three systems examined. Interestingly, the TOC of the intermediates does not significantly depend on the method applied. This phenomenon may be attributed to that in the case of the individual methods (i.e., O_3/UV and $\text{air}/\text{TiO}_2/\text{UV}$) the decay of the mineralization of the intermediates formed is slower than in the combined procedure, where also their formation rate is higher.

Insert Fig. 4 about here

The change of the absorption spectrum of the system during the degradation is in accordance with the formation and decay of intermediates. In the first 50-min period of irradiation the maximum absorbance at about 270 nm did not significantly changed, while a shoulder arose at 290 nm, indicating the hydroxylation of the starting material (Fig. 4A). At

longer periods of irradiation a strong decrease of the absorbance at the main bands can be observed (Fig. 4B), due to the cleavage of the aromatic ring, which is the key step of the decay of the hydroxylated intermediates. The formations of the hydroxylated and ring-opened intermediates in this system are supported by HPLC-MS measurements in our earlier work on the degradation of benzenesulfonate (Szabó-Bárdos et al. 2011). The strong correlations between the absorption spectra and the corresponding MS results in that study unambiguously confirm our interpretation.

Similarly to the case of the TOC values, also the sulfur concentrations belonging to the intermediates (in forms of sulfate or sulfonate groups) could be calculated as functions of the irradiation time (Fig. 5). During the calculation, the total (starting) sulfur concentration was reduced by the actual (free) $[\text{SO}_4^{2-}]$ and by the sulfur content of the residual starting material (pollutant).

Insert Fig. 5 about here

The plots in Fig. 5 indicate that at each point of the degradation process sulfur concentration belonging to the intermediates are highest in the combined procedure. According to these results, the relative sulfur content is highest in the intermediates existing in the system treated by the $\text{O}_3/\text{TiO}_2/\text{UV}$ technique because the concentration of the intermediates during the irradiation is practically independent of the method applied (see Fig. 3). This phenomenon suggests, in accordance with the ratios of the rates of mineralization (TOC decrease, Fig. 1A) and sulfate formation (Fig. 2), that mineralization of the intermediates (i.e. their transformation into carbon dioxide) in the combined procedure is faster than their desulfonation. HPLC-MS monitoring of the intermediates formed in the photocatalytic degradation of benzenesulfonate confirmed that mineralization of these species could take place to a considerable extent without desulfonation (Szabó-Bárdos et al. 2011).

3.4. Mineralization of 2,5-dihydroxybenzenesulfonate (DHBS) under various conditions

Similarly to 4-hydroxybenzenesulfonate (HBS), the mineralization of 2,5-dihydroxybenzenesulfonate (DHBS), the other key intermediate of the degradation of benzenesulfonate, was also investigated by application of the individual methods (O_3/UV and $air/TiO_2/UV$) as well as their combination ($O_3/TiO_2/UV$).

Insert Fig. 6 about here

Comparing the TOC changes during the degradation of DHBS (Fig. 6) to those in the mineralization of HBS, the rates achieved by the corresponding methods are higher for the dihydroxy than for the hydroxy derivative. These results suggest that an enhanced „pre-hydroxylation” of the parent surfactant (benzenesulfonate) promotes the mineralization process. A higher number of hydroxyl groups on the benzene ring increases the probability of the ring-cleavage under the conditions applied in this study. This effect is most pronounced in the case of ozonation, thus, in the mineralization of DHBS by the combined procedure no or just very slight synergistic effect could be observed.

Deviating from the case of HBS, where the absorption change indicated the formation of intermediates (primarily a dihydroxy derivative as the position of the arising new band suggested) in the first period of irradiation (Fig. 4A), a continuous decrease of the absorbance at the main band (above 290 nm) could be observed from the very beginning of the irradiation (Fig. S2). A very slight red shift of the main band indicates the formation of aromatic intermediates (i.e., tri- and tetrahydroxy derivatives) but their decay via ring-cleavage is much faster than their formation, resulting in an overall decrease of the absorbance in this range of wavelength. This observation is in accordance with the enhanced mineralization rates for DHBS compared to those for HBS. The recently published correlations between the spectral

features and the MS values of the corresponding (multihydroxylated aromatic and ring-opened) intermediates (Szabó-Bárdos et al. 2011) also support these results.

3.5. Degradation of alkylbenzene sulfonates (LAS) in aerated and ozonated systems

After the examination of the mineralization of HBS and DHBS, two groups of surfactants frequently applied in industrial technologies were studied in the respect of photocatalytic degradation in both aerated and ozonated systems. The first group involved linear alkylbenzene sulfonates with 10-13 carbon atoms in the alkyl chain. Although these widely used detergents are relatively biodegradable, their biomineralization can be accelerated by photocatalytic pre-treatment.

In the first stage of photocatalytic degradation of these anionic surfactants, air-bubbling could not be used for introducing oxygen (and ozone) as electron acceptor in this system because of the strong foaming. Instead, addition of hydrogen peroxide was applied for this purpose. In this case the reaction mixture was circulated by a liquid (peristaltic) pump. In the first stage, during 24-hour irradiation, 110 cm³ 35% H₂O₂ solution was gradually added to the reaction mixture of 3 dm³.

Insert Fig.7 about here

Fig. 7A displays the decay of the surfactants in two runs of irradiation in the first stage. The two plots clearly indicate the high reproducibility of our results. According to them, within the 24-hour period of irradiation the concentration of the starting surfactants decreased below the foaming limit, practically to zero.

After transformation of the starting detergents in the first stage, air-bubbling could be used for introduction of oxygen (and ozone). In this second stage, irradiation was continued for a 21-hour period under aerated or ozonated conditions. Fig. 7B displays the change of the TOC in the system during both the first and the second stages (the dashed line indicates the

switch between them). Due to the same conditions, in the first stage the plots are very similar, indicating a linear decrease of TOC in the time period of 9-24 h. In the second stage of the mineralization, under aerated conditions the slope of the plot practically did not change compared to that in the first stage, while in the ozonated system the mineralization rate dramatically increased. These results clearly demonstrate that combination of heterogeneous photocatalysis with ozonation considerably enhances (i.e., multiplies) the efficiency of the mineralization of these surfactants.

3.6. Degradation of alkyl ether sulfates (LES) in aerated and ozonated systems

The second group of industrial surfactants examined involved linear alkyl ether sulfates with 10-13 carbon atoms in the alkyl chain. The same types of experiments were carried out with them as with the alkylbenzene sulfonates (see the previous section). In order to avoid foaming, also for these surfactants two-stage procedures were applied. In the first stage, without air-bubbling, hydrogen peroxide was used as oxidizer in a closed reactor with circulation by a liquid pump.

Insert Fig. 8 about here

As Fig. 8A shows, the decay of the LES compounds was much faster in the first stage as that of the LAS surfactants; the concentration of the previous detergents decreased from 520 mg dm⁻³ to 34 mg dm⁻³ (below the foaming limit) within 3 hours, although only 20 cm³ 35% H₂O₂ solution was added to the reaction mixture. The initial rate of the decay of LES was 224 mg dm⁻³ h⁻¹ compared to that of the LAS (56 mg dm⁻³ h⁻¹, see Fig. 7A). This fourfold increase of the decay rate can be attributed to the structural difference between the two types of surfactants. In the case of LES the oxygen-containing ether parts promote the oxidative fractionation of these detergents, while in the LAS compounds no oxygen atoms are involved in the organic parts of the molecules.

Also the rate of the TOC change, both in the first and in the second stage (Fig. 8B) was significantly higher for the LES compounds than for the LAS detergent. This considerable difference ($14.5 \text{ mg dm}^{-3} \text{ h}^{-1}$ vs. $9.6 \text{ mg dm}^{-3} \text{ h}^{-1}$ initial rates in the ozonated systems) can be accounted for the above mentioned structural feature. Moreover, the aromatic ring in the LAS surfactants further hinders their mineralization, due to its extra stability. Hence, 45-hour irradiation was needed for about 85% mineralization of the LAS compounds, while almost total mineralization of the LES compounds was achieved in about 30 hours by the combined procedure ($\text{O}_3/\text{TiO}_2/\text{UV}$). The difference between the efficiencies in the aerated systems is even more pronounced.

The possibility of the reutilization of the TiO_2 catalyst was also tested. After separation by sedimentation and filtration, the used titanium dioxide was rinsed by distilled water, and subsequently applied again under the same conditions as before. In accordance with our earlier observation in a similar system (Szabó-Bárdos et al. 2008a), the activity of the catalyst did not change, i.e., the degradation rate of the pollutants studied did not decreased at all compared to the usage of fresh TiO_2 . This result also supports the practical applicability of the method investigated.

4. Conclusions

Our results clearly demonstrate that the combination of titanium dioxide based heterogeneous photocatalysis with ozonation results in a synergistic effect in the oxidative degradation of hydroxy and dihydroxybenzenesulfonates, achieving a significantly higher mineralization efficiency than the sum of those of the individually applied methods. Manifestation of the advantage of this combined procedure could also be realized by the treatment of widely used industrial surfactants, linear alkylbenzene sulfonates and alkyl ether sulfates. Hence, this combination may be applied for the removal of these pollutants from real wastewaters, either

as a preparatory procedure before the biological treatment or as a final stage, depending on the requirements of purity and costs, and on the type and concentration of the contaminants to be removed

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Figure captions

Fig. 1 Change of the total organic carbon (TOC) content (A) and the concentration of the starting material (HBS) (B) as functions of time in the system containing 1×10^{-3} M HBS during the treatment by various oxidation methods: (\blacktriangle) $O_3 + UV$, (O) air+ $TiO_2 + UV$, (\bullet) $O_3+TiO_2 + UV$. (TiO_2 content is 1 g dm^{-3})

Fig. 2 Change of the sulfate concentration as functions of time in the system containing 1×10^3 M HBS during the treatment by various oxidation methods: (\blacktriangle) $O_3 + UV$, (O) air+ $TiO_2 + UV$, (\bullet) $O_3+TiO_2 + UV$. (TiO_2 content is 1 g dm^{-3})

Fig. 3 Change of TOC belonging to the intermediates during the mineralization in the system containing 1×10^{-3} M HBS and treated by various oxidation methods: (\blacktriangle) $O_3 + UV$, (O) air+ $TiO_2 + UV$, (\bullet) $O_3+TiO_2 + UV$. (TiO_2 content is 1 g dm^{-3})

Fig. 4 Change of the absorption spectrum of the bulk solution during the combined ($O_3+TiO_2 + UV$) treatment of the system containing 1×10^{-3} M HBS and 1 g dm^{-3} TiO_2 . ($\ell = 1.0 \text{ cm}$). The samples were taken at the indicated times in min

Fig. 5 Change of the sulfur content belonging to the intermediates during the mineralization in the system containing 1×10^{-3} M HBS and treated by various oxidation methods: (\blacktriangle) $O_3 + UV$, (O) air+ $TiO_2 + UV$, (\bullet) $O_3+TiO_2 + UV$. (TiO_2 content is 1 g dm^{-3})

Fig. 6 Change of the total organic carbon (TOC) content as functions of time in the system containing 1×10^{-3} M DHBS during the treatment by various oxidation methods: (\blacktriangle) $O_3 + UV$, (O) air+ $TiO_2 + UV$, (\bullet) $O_3+TiO_2 + UV$. (TiO_2 content is 1 g dm^{-3})

Fig. 7 Change of the concentration of the starting material (LAS) during the photocatalytic treatment in the first stage ($H_2O_2 + TiO_2 + UV$) (A); and the total organic carbon (TOC) content in the first ($H_2O_2 + TiO_2 + UV$) and the second stage as well (the dashed line indicates the switch between the two stages) (B) as functions of time in the system containing 1 g dm^{-3} LAS sample and 1 g dm^{-3} TiO_2 . In part A the different markers designate the data points belonging to two independent run of experiments. In the second stage (in part B) two oxidation methods were applied: (O) air+ $TiO_2 + UV$, (\bullet) $O_3+TiO_2 + UV$

Fig. 8 Change of the concentration of the starting material (LES) during the photocatalytic treatment in the first stage ($H_2O_2 + TiO_2 + UV$) (A); and the total organic carbon (TOC) content in the first ($H_2O_2 + TiO_2 + UV$) and the second stage as well (the dashed line indicates the switch between the two stages) (B) as functions of time in the system containing 1 g dm^{-3} LAS sample and 1 g dm^{-3} TiO_2 . In the second stage (in part B) two oxidation methods were applied: (O) air+ $TiO_2 + UV$, (\bullet) $O_3+TiO_2 + UV$

Fig1

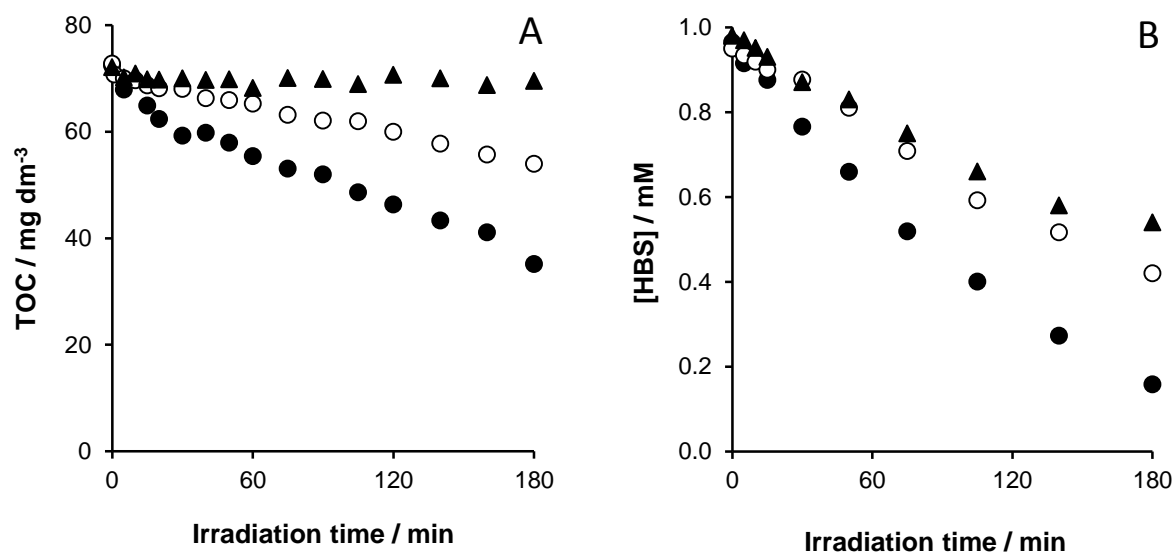


Fig2

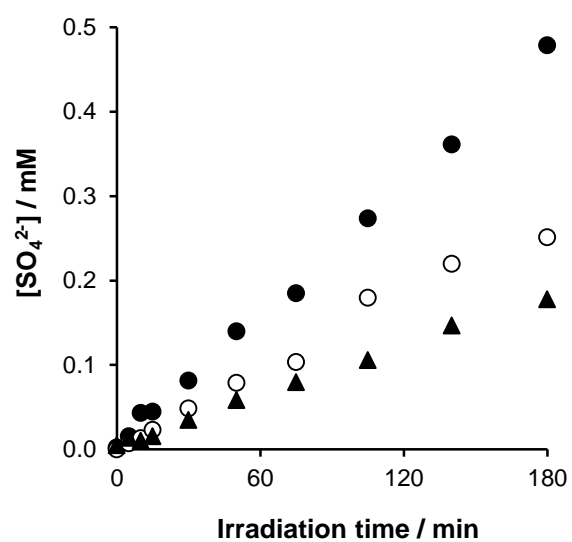


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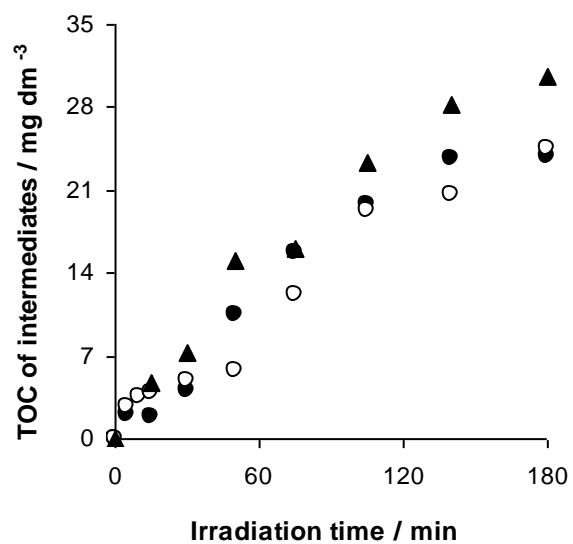


Fig4

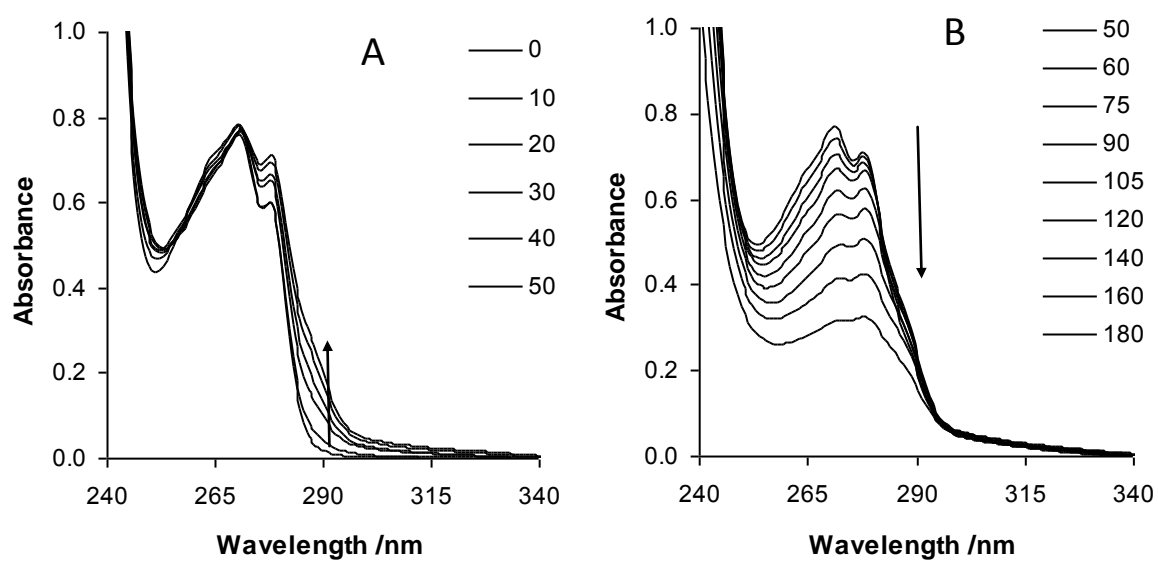


Fig5

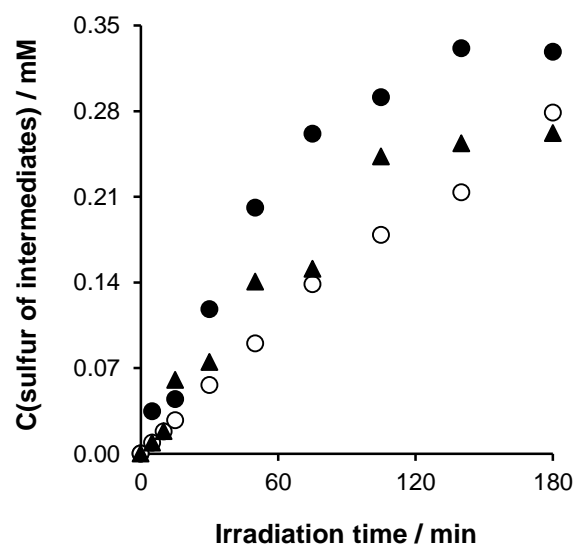


Fig6

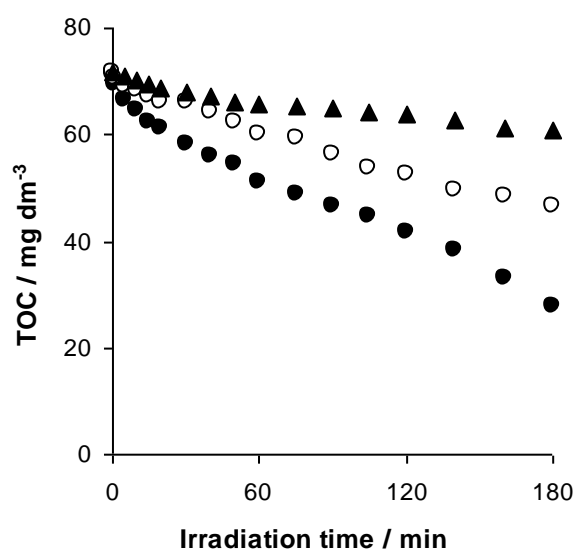


Fig7

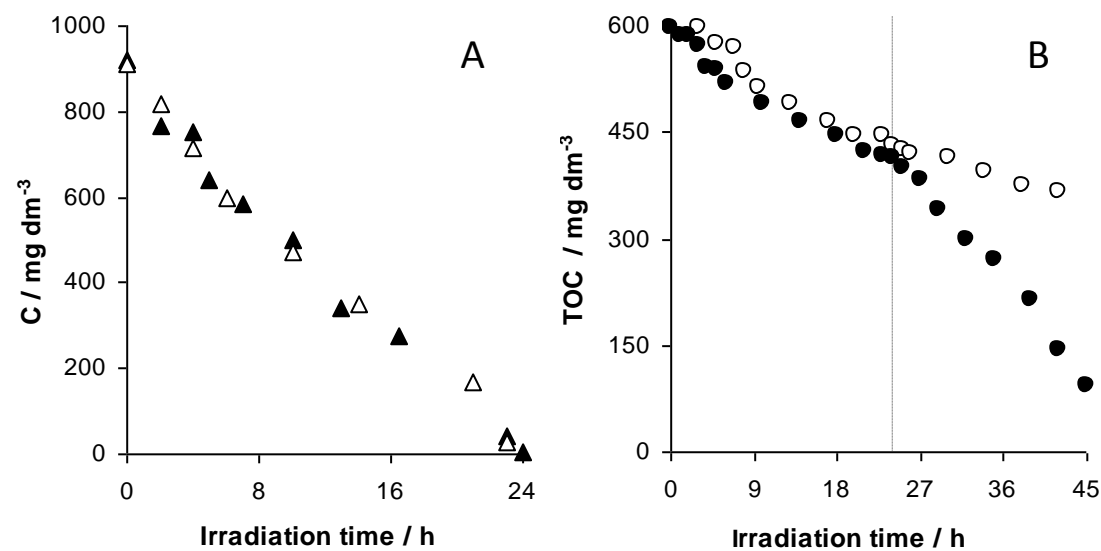


Fig8

