

Spontaneous doping of 2D MoS₂ basal plane by oxygen substitution during ambient exposure

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The chemistry of atomically thin crystals can markedly differ from the conventional surface reactions of their bulk counterparts. Here we provide atomic level insights into the ambient oxidation reaction of MoS₂ monolayers by scanning tunneling microscopy measurements. In contrast to the generally accepted view of environmentally inert basal plane, we found that oxygen atoms spontaneously incorporate into the basal plane of MoS₂ single layers during ambient exposure. While the basal plane was found to be stable against commonly investigated O chemisorption and MoO₃ conversion, an oxygen substitution reaction has been revealed, upon which individual sulfur atoms are one by one replaced by oxygen, giving rise to a novel 2D solid solution type MoS_{2-x}O_x crystal. This oxidation process enables the chemical modification of single atomic sites of 2D crystals opening new routes towards their efficient defect engineering. An eloquent example is that the O substitution sites present all over the basal plane substantially increase the catalytic activity of the 2D MoS₂ crystals for electrochemical H₂ evolution reaction.

Layered materials display thickness-dependent properties when approaching the single layer limit. Chemistry is no exception, as evidenced by the oxidation and hydrogenation of graphene^{1,2}. Surface chemistry is a particularly promising approach for engineering the properties of 2D materials, given their fully exposed atomic structure. However, the chemistry of 2D Transition Metal Dichalcogenide (TMDC) crystals is mainly defined by their edges³, where coordinatively unsaturated sites prevail. Not coincidentally, these reactive edge sites were also found to be responsible for the catalytic activity of MoS₂^{4,5}, one of the most widely studied TMDC materials. However, the chemical modification restricted to edges is only efficient in nanoscale islands⁶, as the edge to surface ratio drastically decreases in the 2D limit. Consequently, for 2D crystals it is of particular importance to chemically and catalytically activate their entire basal plane. Oxidation has widely been studied for graphene as a promising approach towards its efficient exfoliation⁷. For 2D TMDC crystals the oxidation reaction is also of particular importance, as for some crystals it can spontaneously proceed under ambient conditions. Therefore, its study is essential for understanding their long-term environmental stability, as well as it can open new routes towards chemically engineering their properties.

While some TMDC crystals, such as HfSe₂, MoTe₂ or WTe₂, are known to be air-sensitive^{8,9,10} as they rapidly degrade under ambient, the most widely investigated members of the TMDC family (MoS₂, MoSe₂, WS₂, WSe₂) have generally been considered air-stable^{11,12}, based also on decades-long experience with their bulk crystals. Nevertheless, it has been demonstrated that in single layer form the oxidation of MoS₂ and WS₂ also occurs under ambient conditions¹³. The detailed investigations revealed that the oxidation induced etching originates from edges and grain boundaries and proceeds towards the interior of the flakes. Due to the higher strength of the Mo-O bonds, as compared to Mo-S, the substitutional oxidation of the 2D MoS₂ basal plane is in principle also thermodynamically favorable^{14,15}.

However, while at under-coordinated atomic sites on edges and grain boundaries, such oxidation is a fast, low-barrier process^{14,15}, the oxidation of the defect-free basal plane has been predicted to face relatively high kinetic barriers of about 1.6 eV¹⁶, rendering the basal plane environmentally stable. Although experiments often reveal significantly lower activation energy values for the MoS₂ oxidation, such as 0.54 eV or 0.98 eV^{17,18}, no direct evidence on the ambient oxidation of the pristine MoS₂ basal plane has been reported so far. We propose that this is mainly due to the inability of the employed structural characterization methods to resolve single-atom level structural changes. Harsh oxidation processes, such as oxygen plasma treatment, UV-ozone exposure, electrochemical exfoliation, or high temperature (> 300°C) annealing are able to oxidize the basal plane of MoS₂ crystals. Based primarily on XPS investigations it was shown that such processes can lead either to covalent oxygen bonding to the top sulfur atoms^{14,19}, or the formation of completely oxidized MoO₃ areas^{20,21,22} that can subsequently volatilize, leading to etching. However, neither process is ideal for chemically tuning the properties of MoS₂ sheets. While the formation of MoO₃ destroys the original MoS₂ crystal lattice, yielding an overall disordered and fragmented structure, the chemisorption of oxygen onto chalcogen atoms is predicted to have a relatively weak influence on band structure and properties^{14,16}. It has also been shown that oxidation can both enhance and degrade the catalytic activity of MoS₂, depending on the structural details^{23,24}, further emphasizing the importance of investigating and controlling the resulting oxidized structure. The possibility of a substitutional oxidation reaction of MoS₂ has also been raised as being thermodynamically preferred to oxygen chemisorption^{15,25}; however, no clear experimental evidence has been reported so far. The controlled substitutional oxidation of the MoS₂ basal plane would be a highly desirable reaction that preserves the original MoS₂ crystal structure, while in contrast to O chemisorption, it is also expected to substantially

influence the electronic band structure, enabling a more efficient engineering of its electronic^{14,16} and optical²⁶ properties.

Here we show that the basal plane of the MoS₂ monolayers subjected to long-term ambient exposure spontaneously undergo such oxygen substitution reactions giving rise to a novel, highly crystalline two-dimensional molybdenum oxy-sulfide phase.

Results and discussion

Ambient oxidation of the basal plane revealed at single-atom level

We have prepared mechanically exfoliated MoS₂ single layers on atomically flat Au (111) substrates using a slightly modified version of a recently developed exfoliation technique²⁷ yielding single layers with lateral dimension of hundreds of microns (see SI section I. for details). Such large area samples are characterized by an extremely low edge to surface ratio, grain boundary concentration, as well as a much smaller concentration and variety of intrinsic point defects²⁸, establishing them as an excellent model system for studying the intrinsic chemical properties of the pristine basal plane. The exfoliated MoS₂ samples have been stored under ambient conditions (air, room temperature and ambient light) for periods of up to 1.5 years. We have employed STM measurements to follow the atomic level structural changes in the basal plane structure of 2D MoS₂ crystals during long-term ambient exposure. So far mainly optical, scanning electron and atomic force microscopy measurements have been employed to monitor the structural changes induced by oxidation in MoS₂ layers. However, the spatial resolution of these methods does not allow detecting single atom level modifications. Imaging the atomic-scale structure of the MoS₂ basal plane is possible by high resolution scanning Transmission Electron Microscopy (TEM). However, detecting light atoms such as oxygen with TEM is challenging due to their low contrast and easy knock-out^{29,30}. By contrast, STM measurements can detect the unaltered structure of such oxygen defects due to its atomic resolution capability and the low energy of the

tunneling electrons³¹. Due to the slow oxidation reaction, and the noninvasive nature of the measurements, the structure and number of oxidation induced defect sites did not change during the STM imaging, enabling us to acquire atomic resolution snapshots of the oxidation process.

Atomic resolution STM images of the basal plane of a mechanically exfoliated MoS₂ single layer after 1 month and 1 year of ambient exposure are shown in Fig. 1b and 1c, respectively. The STM images have been acquired on the same sample, but at different locations. In contrast to general expectations, STM measurements reveal clear modifications in the atomic structure of the MoS₂ basal plane during ambient exposure. Freshly prepared 2D MoS₂ crystals contain a native point defect density in the range of 10^{11} - 10^{12} cm⁻² (Fig. S2). After a month of ambient exposure our STM measurements revealed the formation of new point defects all over the basal plane increasing their concentration into the 3×10^{12} - 2×10^{13} cm⁻² range (Fig. 1b). The increase of the atomic scale defect concentration (up to 5×10^{13} - 10^{14} cm⁻²) appears more strikingly after a year exposure (Fig. 1c), when a substantial area of the sample surface is already covered by such defects, clearly evidencing the ability of the ambient exposure to create new defect sites in the basal plane structure of 2D MoS₂ crystals. Although the defect density shows some spatial variation, as evidenced by the concentration ranges provided above, the effect of the ambient exposure time is more significant. Higher resolution STM images (Fig. 1d) shed light on the atomic structure of the defects induced by ambient exposure. Dark triangles in the STM images of 2D MoS₂ crystals are characteristic to sulfur atom vacancies³¹. Within the dark triangles, bright spots can be clearly detected that have not been observed before. These central bright spots are a pronounced feature of the experimental STM images, as their contrast is even stronger than that of S atoms of the MoS₂ lattice. The most straightforward choice is to attribute the bright spots inside the S vacancies to atoms or molecules saturating the vacancies under ambient

conditions. Among ambient species oxygen has been predicted to be the energetically most favorable to saturate an S vacancy^{14,16}. To confirm this, we have performed detailed simulations of the STM images of various O-related defects in 2D MoS₂ crystals based on DFT calculations of their electronic structure (Fig. S3). Our theoretical results reveal that the best agreement with the experimental STM data is provided by single oxygen atoms substituting individual S atoms (saturating the S vacancy) displaying a dark triangle with a bright spot inside. Such a simulated STM image is shown in the inset of Fig.1d. We note that according to both calculations and measurements the contrast of the O substitutions is dependent on the tip-sample distance; it only becomes fully apparent at small tip-sample distances (see SI section IV for details). While the agreement with experiments is far better than for any other calculated defect structure, it is not perfect. This can be attributed partly to the STM image distortions that are unavoidable at room temperature, and partly to the dark ‘halo’ often emerging near single atomic substitution sites in STM images^{32,33}. The latter can be a signature of an electron depletion zone around a negatively charged defect site³⁴. STM investigations cannot directly confirm the chemical identity of the incorporated atoms, XPS measurements are usually employed for such purposes. However, due to their low spatial resolution, it was not possible to selectively measure the single layer areas, as exfoliated samples are not homogeneously covering the substrate. Furthermore, the relatively low concentration of the incorporated oxygen atoms (< 3 at%) makes such investigations challenging¹³. To gain more information on the nature of the defects progressively forming under ambient conditions, we have also performed Raman spectroscopy (Fig S5) and photoluminescence (PL) measurements. For the PL measurements we have transferred the MoS₂ single layers from Au (111) to SiO₂/Si substrates. PL spectra on aged samples revealed a new peak at 1.75 eV that is not present in freshly exfoliated samples (see Fig S6). This peak can be clearly associated with vacancy type defect formation^{35,36}. Moreover, it has been

shown that such peak does not emerge under ultra-high vacuum conditions; the charge transfer induced by the saturation of vacancies with environmental species (oxygen, nitrogen) is necessary³⁶. These findings further support our STM and simulation data on the formation of oxygen saturated S vacancies.

The above findings clearly challenge the generally accepted view of environmentally inert MoS₂ basal plane, and evidence that the ambient oxidation of the basal plane spontaneously proceeds through formation of S vacancies and their saturation by oxygen. This oxidation process preserves the original crystal lattice with Mo sites in the trigonal prismatic configuration, but coordinated by five sulfur and an oxygen atom. The oxidation speed of the MoS₂ basal plane under ambient was found to be of order of 1 atom / minute / μm^2 . This ultra-slow oxidation reaction in principle enables an extremely precise control of oxygen concentration in the MoS₂ lattice. The slow reaction speed is most likely responsible for stabilizing the oxy-sulfide phase against the otherwise favorable full conversion to MoO₃²⁰.

Since the oxidation process is expected to be kinetically limited^{14,16}, by increasing the temperature, the oxidation speed should also increase. We have subjected a freshly exfoliated MoS₂ /Au (111) sample to air inside a furnace heated to 400 K for one week. Subsequent STM investigation revealed a similar defect formation as at room temperature, but the oxygen site concentration reached values of $8 \times 10^{12} - 3 \times 10^{13} \text{ cm}^{-2}$ already after a week of exposure (see Fig. S7). This clearly indicates that the oxygen substitution process has been accelerated by increasing the temperature. Furthermore, this observation also excludes the possibility that the ambient oxidation occurs due to the accidental presence of some more reactive oxygen species (e.g., ozone, superoxide, singlet oxygen). Increasing the temperature from 21°C to 127°C is not expected to significantly increase the formation of such reactive species. The accelerated oxygen substitution process at higher temperatures also provides a more feasible route for the synthesis of oxy-sulfide crystals.

The substitutional oxidation of MoS₂ basal plane yields a novel 2D MoS_{2-x}O_x solid solution crystal phase (Fig. 1a). The formation of MoS_{2-x}O_x solid solution phase has already been proposed for sputter deposited MoS₂ films^{37,38,39}. However, so far the structure of the synthesized molybdenum oxy-sulfide films was found to be amorphous or highly disordered with poor long-range crystalline order^{40,41}. Here we provide clear evidence for the formation of a highly crystalline 2D molybdenum oxy-sulfide phase. This provides a novel strategy for synthesizing new materials through chemical transformation of 2D crystals that are difficult to synthesize by other methods.

The STM investigations of 2D MoSe₂ crystals (prepared and aged under the very same condition as 2D MoS₂ crystals) revealed a strikingly different behavior upon long-term ambient exposure. Fig. 2 shows atomic resolution STM images of freshly exfoliated and 1 year old MoSe₂ single layers. The basal plane of the 2D MoSe₂ crystal does not show significant changes even after year-long ambient exposure. Native defects (Se vacancies and chemisorbed adatoms – see the inset of Fig. 2a) are present with a relatively low (10^{11} - 10^{12} cm⁻²) density, but their concentration is not increasing during ambient exposure.

The energetics and kinetics of oxygen substitution

Based on the existing literature, it is difficult to interpret the experimentally observed striking differences in the ambient oxidation behavior of 2D MoS₂ and MoSe₂ basal planes. However, most theoretical works so far focused on the chemisorption mechanism. Indeed, our experimental results confirm that the pristine basal planes of both MoS₂ and MoSe₂ single layers are stable against chemisorption or MoO₃ transformation. We found that oxygen substitution proceeds in the case of MoS₂, but not for MoSe₂. In contrast to chemisorption, the oxygen substitution mechanism implies the removal of S (Se) atoms during the oxidation process. To investigate the thermodynamics of this process, we have performed DFT calculations (see Methods for details) regarding the removal of S (Se) atoms from the basal

plane of the 2D MoS₂ (MoSe₂) crystals through oxidation. First, we have investigated the reaction of surface sulfur atoms of 2D MoS₂ crystals with oxygen, yielding volatile SO₂ species and creating a sulfur vacancy in the MoS₂ basal plane (Fig. 3a). The enthalpy of the reaction is calculated using the formula: $\Delta E = E(\text{Reactants}) - E(\text{Products}) = -0.49 \text{ eV}$. Since the enthalpy of the final structures is lower than the initial, this implies that S atom removal by oxidation is thermodynamically favorable in the case of MoS₂. By contrast, our DFT calculations show that the removal of a Se atom through oxidation of MoSe₂ basal plane displays a positive oxidation enthalpy of $\Delta E = +0.75 \text{ eV}$ (Fig. 3b). This result evidences the thermodynamically unfavorable nature of selenium oxide formation, which hinders the formation of Se vacancies through oxidation and hence the substitutional oxidation of the MoSe₂ basal plane, in agreement with the experimental observations.

Even though thermodynamically favorable, the substitutional oxidation of MoS₂ basal plane is still expected to face kinetic barriers. To determine this we have performed detailed nudged elastic band (NEB) model calculations⁴² for determining the energy of transitional states and finding the potential barriers for the substitutional oxidation process. The results displayed in Fig 3c show that in the case of MoS₂ the typical kinetic barrier height is about 1 eV for the proposed reaction pathway. We note that by considering the O saturation of the S vacancies the energy of the final state becomes by about -4 eV lower (Fig S3), indicating the highly favorable nature of the O substitution process as a next step, in accordance with the experimental findings. By contrast, the barriers for the substitutional oxidation of MoSe₂ basal plane are about 1.5 eV, with a thermodynamically unfavorable final state (Fig. 3d).

Consequently, both energetics and kinetics support our experimental findings on the differences between the oxidation of the 2D MoS₂ and MoSe₂ basal planes, reflecting the different oxidation mechanisms of sulfur and selenium based compounds. Furthermore, the kinetic barriers of about 1 eV height for the substitutional oxidation of MoS₂ are predicted to

be surmounted at room temperature on month's timescale, according to transition state theory⁴³.

Although the basal plane of the MoSe₂ appears more stable, the overall environmental stability of the crystals is also dependent on their edge oxidation speed. We found that the oxidation proceeds much faster at the edges of MoSe₂ single layers than on the MoS₂ edges (Fig. S8, S9). Adsorbed contaminations are naturally present on both MoS₂ and MoSe₂ surfaces subjected to ambient. However, their role in the ambient oxidation process is expected to be more pronounced at edges and grain boundaries^{13,15}, as they are known to preferentially attach to these high-energy sites instead of the pristine basal plane.

While structural disorder induced by invasive oxidation is unlikely to be fully reversible, the substitutional oxidation of the MoS₂ basal plane does not damage the crystal lattice, so it can in principle be suitable for a fully reversible reduction. Indeed, we found that a simple annealing of the MoS_{2-x}O_x crystals in H₂S atmosphere at 200°C for 30 minutes is able to fully restore the atomic structure of the pure 2D MoS₂ phase. Two representative atomic resolution STM images of the MoS₂ basal plane before and after the reduction (Fig. 4) clearly evidence that 2D MoS_{2-x}O_x solid solution crystals can be reduced to the pure, almost defect-free MoS₂ phase. The feasibility of such a reduction process is also supported by our DFT calculations (Fig. S10). The atomically perfect reduction of the oxidized 2D MoS_{2-x}O_x crystals enables a highly efficient engineering of their surface chemistry.

Catalytic activity of 2D MoS_{2-x}O_x crystals towards hydrogen evolution

To investigate how the oxygen substitution sites change the properties of MoS₂ single layers, we have investigated the catalytic activity of 2D MoS_{2-x}O_x crystals for the electrochemical hydrogen evolution reaction (HER) (see Methods for experimental details). Polarization curve (I-E) measurements have been performed on both MoS_{2-x}O_x and MoS₂ single layers. To make the comparison more direct and relevant, we have measured a 1 year

old $\text{MoS}_{2-x}\text{O}_x$ flake before and after its reduction to the pure 2D MoS_2 phase. This way, we could measure the catalytic activity of the very same flake both with and without O substitution sites in the basal plane, as revealed by our STM characterization. The measured polarization curves and the corresponding Tafel plots shown in Fig.5 evidence a highly increased catalytic HER activity of the 2D $\text{MoS}_{2-x}\text{O}_x$ solid solution crystals, as compared to the reduced pure 2D MoS_2 phase. Consequently, this enhanced catalytic activity can be clearly related to the presence of substitutional O sites, the only structural feature that displays a correlation with ambient exposure time, according to atomic level structural STM data. We have also investigated the stability of the 2D $\text{MoS}_{2-x}\text{O}_x$ crystals during 1000 catalytic cycles. After an initial slight decrease by about 10%, the current density stabilizes, indicating a good long-term stability of the catalytic process (Fig. S11). Confocal Raman microscopy maps revealed that the 2D $\text{MoS}_{2-x}\text{O}_x$ layer is still continuous after 1000 cycles and the measured Raman spectra remained practically unaltered (Fig. S12), while atomic resolution STM measurements confirmed that the O substitution sites are present after the catalytic process (Fig. S11).

A widely used parameter for predicting the catalytic activity of various sites is the hydrogen adsorption Gibbs free energy (ΔG_H). We have calculated the ΔG_H for O substitution sites and compared it to that of pristine MoS_2 surface (see SI section X for calculation details). We found that ΔG_H is lowered to about half on the O sites (+ 1.2 eV) as compared to the S sites (+ 2.2 eV) of the basal plane, which means that hydrogen is much more likely absorbed on the O sites. However, even the substantially reduced ΔG_H value of about 1 eV is still well above zero, indicating a not very favorable H adsorption. Nevertheless, other effects - not captured by ΔG_H – can also play an important role in the catalytic activity. Since O atoms have a similar electron configuration to S atoms, it is often assumed that the O substitutions are fully passivating the S vacancies without significantly altering the electronic structure.

While O substitutions indeed remove the midgap states characteristic to sulfur vacancies⁴⁴, they still substantially change the orbital composition of the valence and conduction bands (see SI section XI), as well as the electron affinity at the substitution sites. The electronegativity and electron affinity of the dopants (including O substitutions) has been shown to play an important role in defining the catalytic activity of graphene^{45,46}. To evaluate this effect for the O substitution sites of 2D MoS₂ crystals, we have performed the Bader charge analysis (see SI section XII for details) that has proven useful for understanding the catalytic activity of sites where charge transfer plays an important role⁴⁷. The Bader analysis evidenced a strong acceptor type behavior of the O substitution sites, characterized by almost two-times higher electron affinity (-0.88 e^-) as compared to S atoms (-0.47 e^-). The locally increased electron affinity combined with the experimentally measured overall *n*-doping of the MoS₂ crystals (Fig S15) can give rise to localized negative charges on the O substitution sites. These partially screened negative charges can substantially facilitate the adsorption of positively charged H species from the acidic electrolyte. The effect of such charged dopants is not included in the ΔG_H calculations, as it is challenging to treat charged impurities at the DFT level⁴⁸. Therefore, we propose that the combined effects of the substantially decreased ΔG_H and increased electronegativity at the O sites - both facilitating the absorption of H species - can give rise to the increased catalytic activity observed in 2D MoS_{2-x}O_x crystals.

Understanding the catalytic process is a highly challenging task, in big part due to the complexity of the experimentally investigated systems in strong contrast with the idealized theoretical models⁴⁹. 2D MoS_{2-x}O_x provides an ideal model system for understanding the atomic level relations between active sites and catalytic HER activity, as it is characterized by a single type of active site, with experimentally known atomic structure, while edges and previously investigated more disordered MoS₂ structures can host a variety of active sites with complex atomic configurations and often little experimental insight into their precise atomic

nature. These findings clearly evidence that the substitutional oxidation process of the MoS₂ basal plane reported here can open new routes for engineering 2D electrocatalysts with single oxygen atom active sites of a much higher site density than previously achieved for individual hetero-atom catalysts⁵⁰.

Methods

Sample preparation. The investigated MoS₂ (and MoSe₂) single layers have been prepared by mechanical exfoliation of bulk MoS₂ (MoSe₂) synthetic crystals of high structural quality (2DSemiconductors). We have employed a slightly modified version (see SI section I) of the mechanical exfoliation technique developed by us and discussed in details in ref. (27), providing single layer TMDC flakes with hundreds of microns lateral dimensions on atomically flat Au (111) surfaces. The single layer nature of the investigated MoS₂ and MoSe₂ crystals has also been confirmed by Raman spectroscopy. The samples have been stored under ambient laboratory conditions: in air, at room temperature and ambient light conditions.

Characterization. Atomic resolution STM measurements have been performed on a Nanoscope E STM operating under ambient conditions. Tunneling spectroscopy data have been performed in an RHK PanScan STM under UHV conditions at room temperature. Large area flakes could be easily identified under an optical microscope enabling the guided landing of the STM tip on MoS₂ single layers. Atomic resolution imaging could be achieved under ambient conditions with typical parameter ranges of $|U_{\text{bias}}| = 5 - 50$ mV, and $I_{\text{tunnel}} = 1 - 3$ nA. Although the resolution of the hexagonal lattice of the top layer of S atoms can be routinely achieved, atomic resolution of individual point defects is challenging. Defects were clearly resolved when imaging at bias voltages (energies) within the band gap. This is possible due to the influence of the Au(111) substrate, which induce a small but finite density of states within the band gap (confirmed by our tunneling spectroscopy measurements), conferring a weak metallic character to the supported MoS₂ single layers. Raman measurements have been

conducted on a confocal Raman microscope (Witec 300RSA) with $\lambda = 532$ nm and $W = 1$ mW. We have ensured that during the Raman measurement the spectra of the sample did not change detectably.

Electrochemical measurements. The measurements have been conducted on selected sample areas (0.4 – 0.8 mm diameter) of a single 2D MoS₂ flake supported by a 100 nm thick Au(111) film on a glass substrate. Control measurements on the same Au (111) substrate as well as a Pt plate have been conducted in the same experimental configuration. To compare the performance of different samples in HER, linear sweep voltammetry was performed in a three-electrode configuration using 0.5M sulfuric acid electrolyte at room temperature. Silver / silver chloride and Pt wire were used as counter and reference electrodes, respectively. Potential sweeps were acquired at a scan rate of 2 mV/s using a Bio-Logic SP-150 potentiostat.

Computational details. All calculations were performed on TMDC single layers in the framework of spin-polarized DFT theory implemented in the VASP software package, using the plane wave basis set and projector augmented wave method. Exchange-correlation effects were taken into account in the framework of local density approximation (LDA) and generalized gradient approximation (GGA) by Perdew–Burke–Ernzerhof (PBE) functional. The defects were modelled in periodically repeated 8×8 supercells. The Brillouin zone of the supercells was sampled with (2×2×1) Monkhorst–Pack mesh of k-points for geometry optimization and (4×4×1) for STM-image calculations. The cutoff energy for the plane-wave basis set was set to be 400 eV. To avoid artificial interactions between periodic replicas of low-dimensional nanoclusters, a vacuum interval of 15 Å was introduced in all supercell geometry. The nudged elastic band (NEB) method was applied for modeling transitional states and finding potential barriers of MoS₂ and MoSe₂ oxidation.

Figures

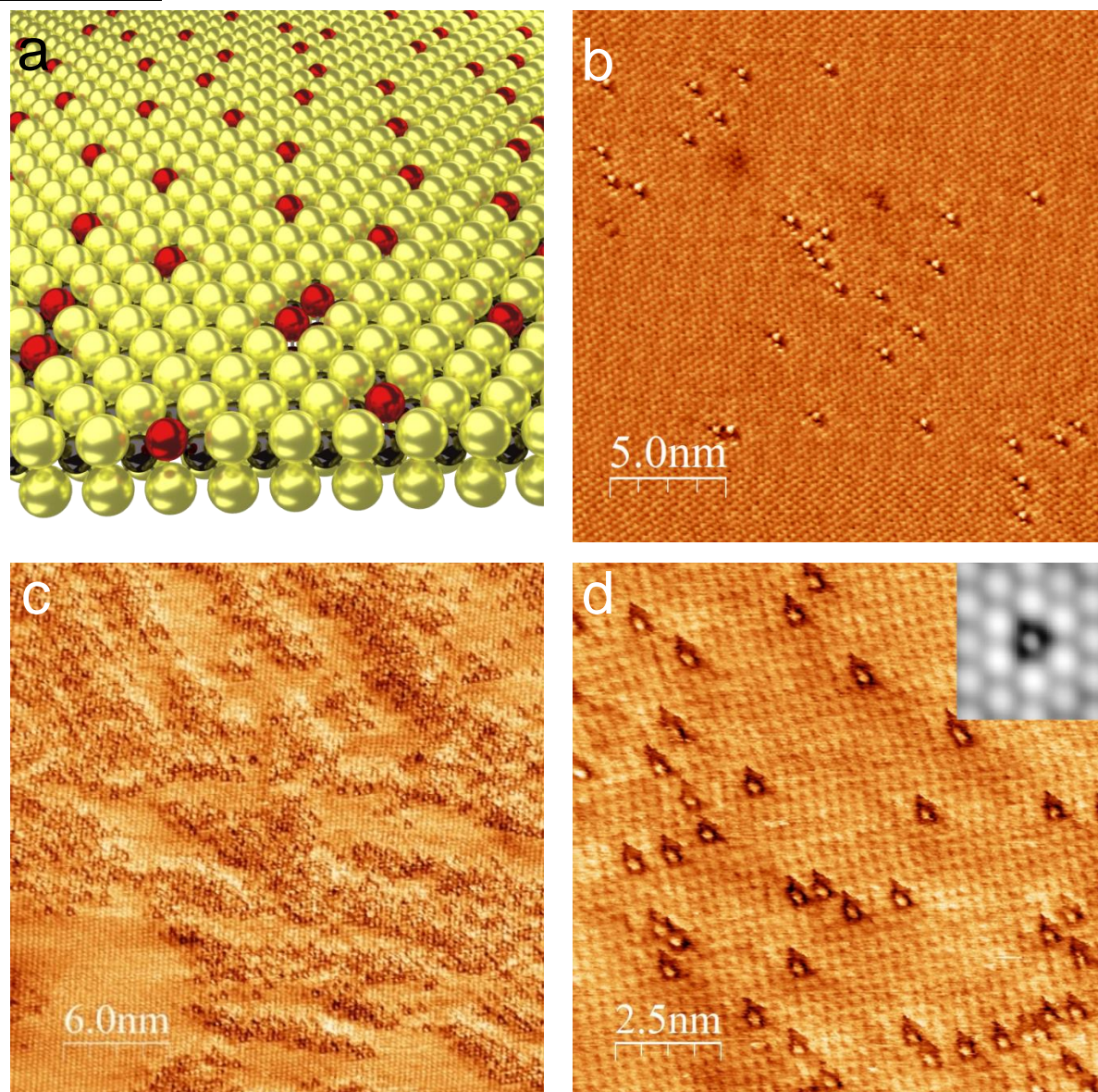


Figure 1. 2D $\text{MoS}_{2-x}\text{O}_x$ solid solution crystals by ambient oxidation of MoS_2 . a) Schematic atomic structure of the $\text{MoS}_{2-x}\text{O}_x$ solid solution monolayer. b)-d) Atomic resolution STM images (5mV, 2nA) of an exfoliated MoS_2 single layer after one month (b) and one year (c) ambient exposure. d) Higher resolution STM image revealing the incorporation of the oxygen (bright spots) into the S vacancies (dark triangles). Inset shows simulated STM image based on DFT calculations of an oxygen saturated S vacancy site in the 2D MoS_2 crystal.

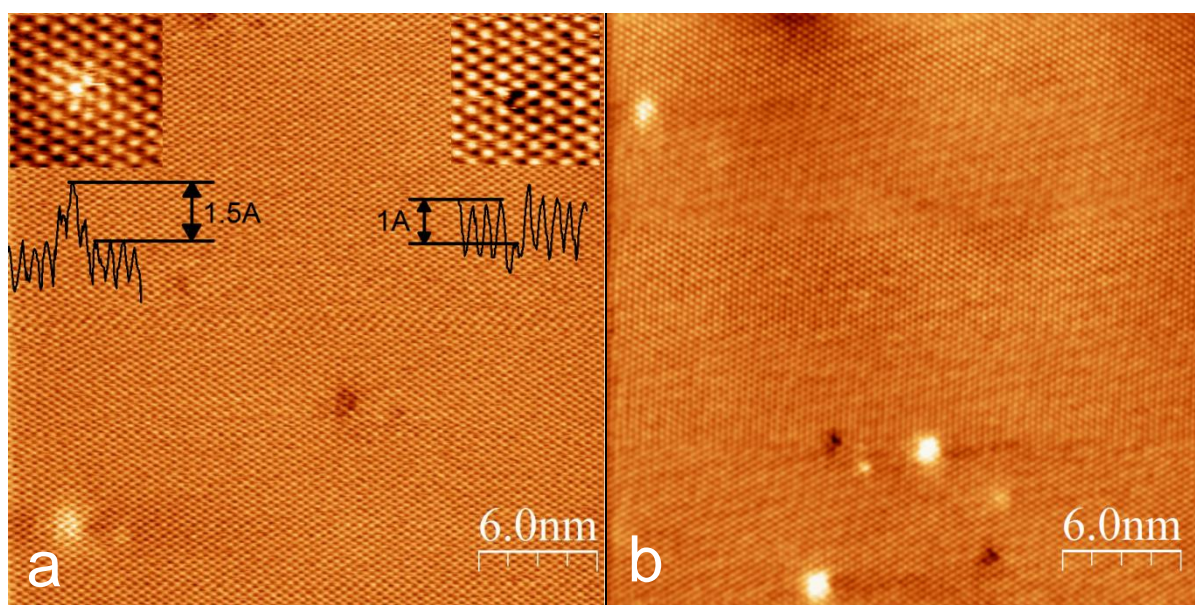


Figure 2. Ambient stability of 2D MoSe₂ basal plane. Atomic resolution STM images (5mV, 1nA) of 2D MoSe₂ crystals on Au (111) substrate freshly prepared (a) and after 1 year of ambient exposure (b), revealing a remarkable ambient stability of the basal plane. The insets of panel (a) show higher resolution STM images and the corresponding line cuts of bright (chemisorbed adatom) and dark (Se vacancy) defect sites.

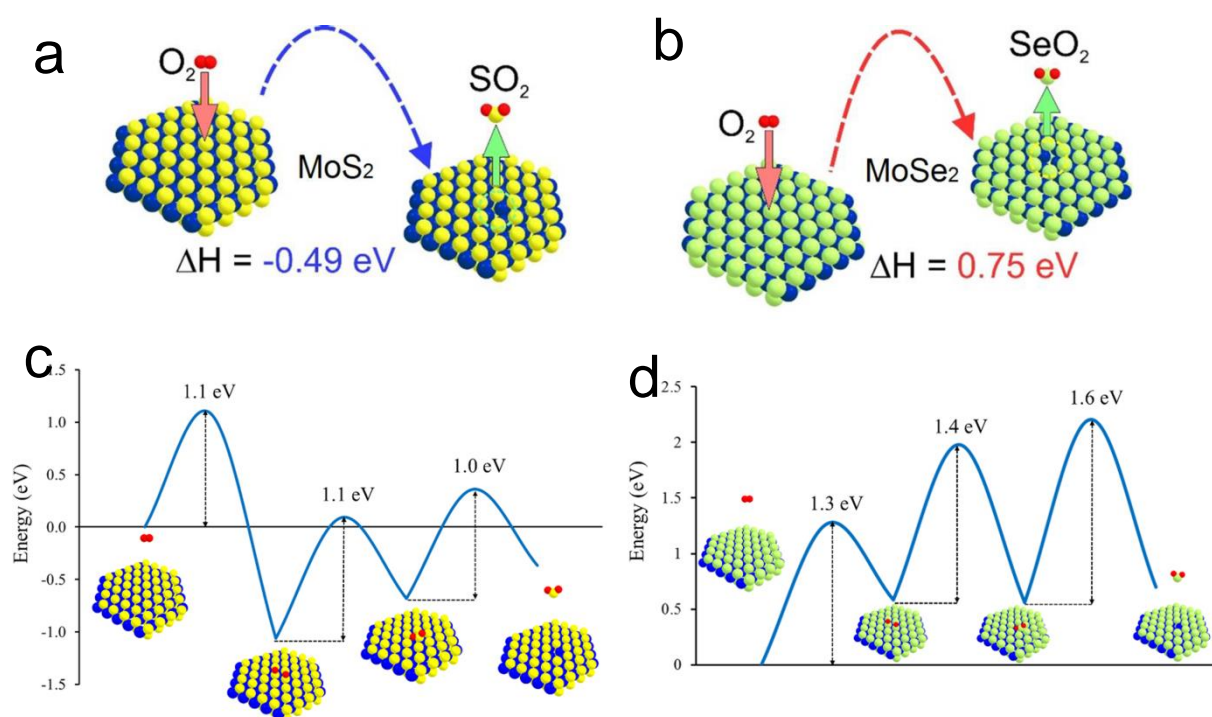


Figure 3. Energetics and kinetics of the oxygen substitution in the MoS₂ and MoSe₂ basal plane. Vacancy formation through oxidation displays a negative oxidation enthalpy for 2D MoS₂ (a) while it is positive for MoSe₂, protected from oxidation by the endothermic nature of SeO₂ formation (b). Kinetic energy barriers calculated by the nudged elastic band model for 2D MoS₂ (c) and 2D MoSe₂ (d) crystals.

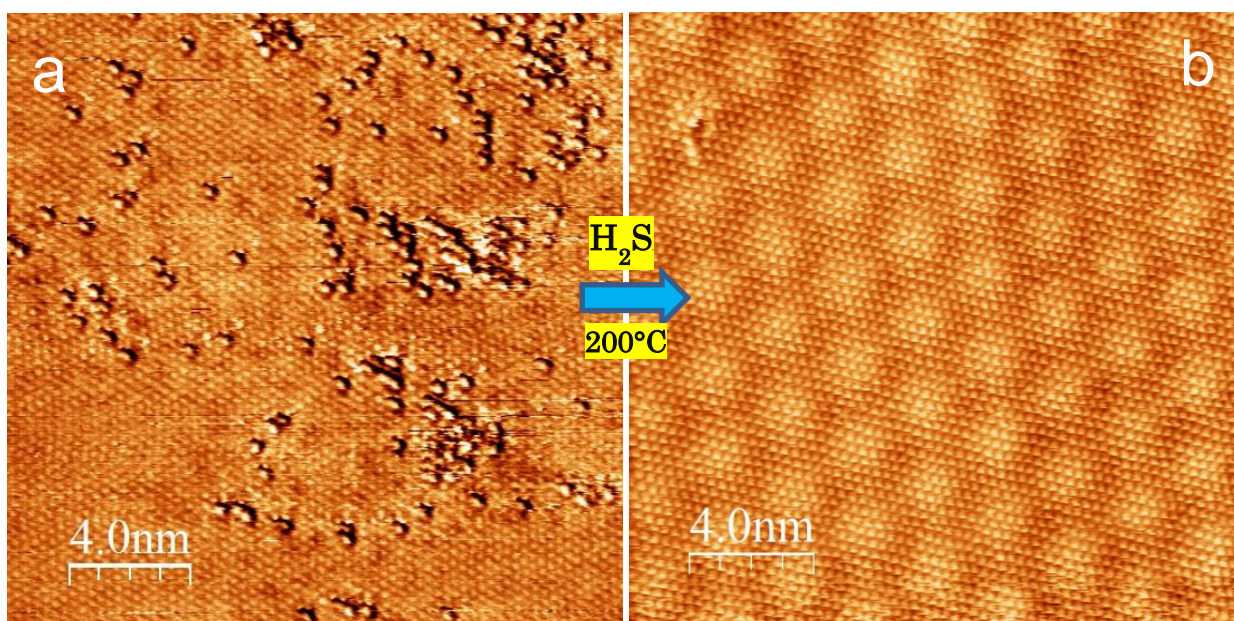


Figure 4. Reduction of 2D $\text{MoS}_{2-x}\text{O}_x$ to perfect MoS_2 . *Representative atomic resolution STM images (5mV, 2nA) of 2D $\text{MoS}_{2-x}\text{O}_x$ before (a) and after (b) 30 minutes annealing at 200° C in H_2S atmosphere, evidencing the atomically perfect reduction of the oxy-sulfide solid solution to the pure MoS_2 phase.*

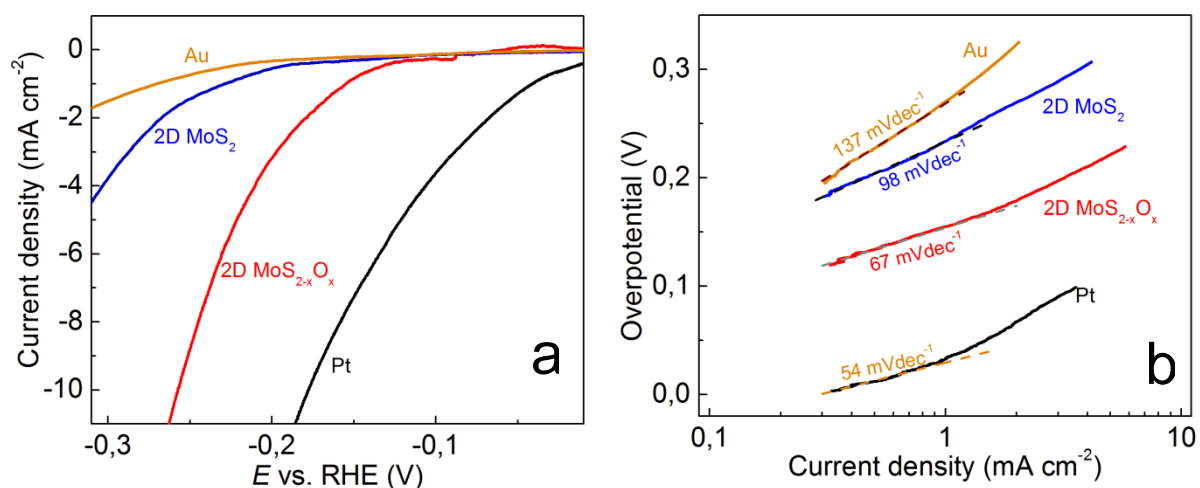


Figure 5. Catalytic activity of 2D MoS_{2-x}O_x for hydrogen evolution. *Linear sweep voltammogram curves (a) and the corresponding Tafel plots (b) for: Au substrate, MoS₂ single layer, MoS_{2-x}O_x single layer (1 year old), and Pt substrate, revealing a significantly higher catalytic activity of the 2D oxy-sulfide phase as compared to the pure MoS₂ phase, attributed to the novel catalytically active O substitution sites emerging on the basal plane during the substitutional oxidation process.*

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